Chem 20 Test 5 Review

Tro 3rd Edition

You will be given:

- A periodic table
- Electronegativity chart for polarity
- Conversion factors such as 1000 mL = 1L etc

Memorize

Molarity= moles/Liters $\qquad\qquad$ pH= -log[H_3O^+]

$C_1 \times V_1 = C_2 \times V_2$ $\qquad\qquad$ [H_3O^+][OH^-] = 1 x10^{-14}

Chapter 13 Sections 1-7

1. Define the terms solute, solvent, and solution # 1, 3, 27, 29
2. Define the terms saturated, supersaturated, and unsaturated; miscible and immiscible #6, 33, 35
3. Know the three factors increase solubility- temperature, pressure and "like dissolves like"
4. Given grams and volume, solve for molarity # 61, 63, 65
5. Given molarity and volume, solve for grams # 67, 73, 75, 77
6. Do calculations involving diluting a solution # 83, 85, 87, 89
7. Do calculations involving percent by mass #43, 45
8. Do stoichiometry problems with solutions # 93, 95, 97

Chapter 14 Sec1-6, 8, 9

1. Know the properties of acids and bases # 6, 9
2. Be able to do titration calculations # 53, 55
3. Correctly write neutralization reactions # 13, 45
4. Be able to do calculations with pH, pOH, H_3O^+ and OH^- # 67, 69, 71, 73, 75, 77, 79, 81, 109, 111